

# Determination of the Equivalence Point & NaOH Concentration

تركيب الطحاح الخليوي



1- Prepare 100ml Of [0.1M] of (HCl) from **Lab stock** with these information

Faculty Member (B.Sc. /M.Sc. /D.Sc. Major in "Instrumental Analysis"-Expert & Advanced Major in "Analysis")  
 أساتذة (ب.ع. /م.ع. /د.ع. تخصص في "التحليل الآلي"-الخبرة في "التحليل الآلي"-الخبرة في "التحليل الآلي")  
 FAAS FAES GC GC-MS HPLC IEC ICP-OES ICP-MS ICP-AES FTIR & NMR

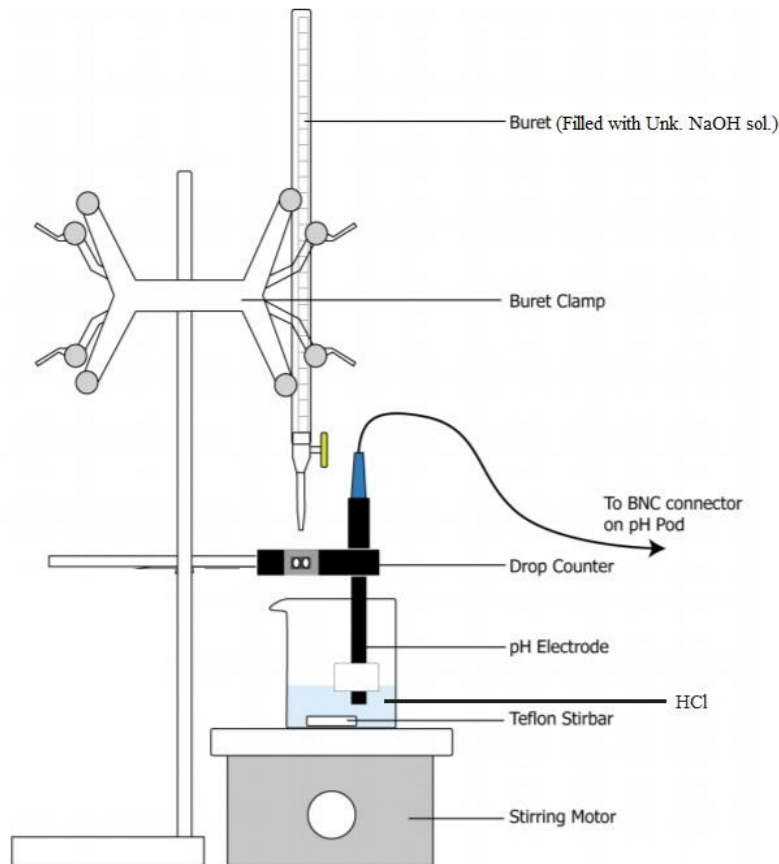
[(Density=1.17g/ml) & (Purity=37% w/w)]. العنوان: كلية العلوم - مبنى (ب) - الدور الأول - الطحاح الخليوي - الرياض 0114670404

2- Take (7.5ml) of solution in [1] in 50ml V.flask and fill with dis.water.

3- Pour [2] in suitable beaker.

4- Fill the burette with (25ml) of given (NaOH) solution.

5- Construct the system as shown in the photo below:



6- Merge the operated pH electrode in the (HCl) beaker and record the pH Value.

7- Drop (1 ml) of (NaOH) from the burette and record the pH Value.

8- Repeat step [7] till you reach the Equivalence Point (1ml each time)..

9- Find the Equivalence Point and calculate the (NaOH) concentration.

## Results:

No	NaOH Added Volume(ml)	pH Value
1	0	$pH_1 < 7$
2	1	$pH_2$
3	2	$pH_3$
4	3	$pH_4$
5	4	$pH_5$
6	5	$pH_6$
7	6	$pH_7$
8	7	$pH_8$
9	8	$pH_9$
...	...	$pH_{...}$
...	...	$pH_{...}$
..n	n	$pH_n > 7$

## Titration Graph:

