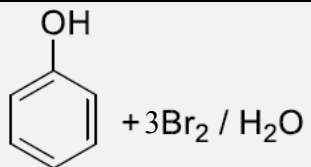


Laboratory Report (109 chem)

Experiment 5: Hydroxy Compounds (Alcohols and Phenols)

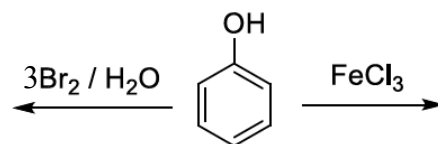
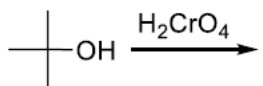
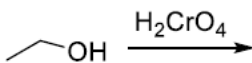
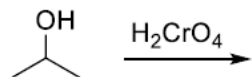
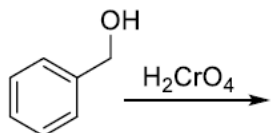
Student Names: Section No:

Test	Observation	Result	Chemical equation
<p style="text-align: center;">1^o alcohol + Conc.HCl</p> <p>1 mL of Ethanol + 1mL of Lucas reagent + heating for 1-2minutes</p>			
<p style="text-align: center;">2^o alcohol + Conc.HCl</p> <p>1 mL of Isopropanol + 1 mL of Lucas reagent + heating for 1-2minutes</p>			
<p style="text-align: center;">3^o alcohol + Conc.HCl</p> <p>1 mL of t-butanol + 1mL of Lucas reagent</p>			
<p style="text-align: center;">1^o alcohol + KMnO₄</p> <p>1mL of Ethanol + 1 drop of KMnO₄ + heating for 1 minute</p>			
<p style="text-align: center;">3^o alcohol + KMnO₄</p> <p>1mL of t-butanol + 1 drop of KMnO₄ + heating for 1 minute</p>			
<div style="text-align: center;">  <p>+ FeCl₃</p> </div>	violet color appears		



Bromine color disappears and a white precipitate forms

Questions:



Name	class	Functional group	Molecular formula	Structure formula
Ethanol				
Isopropanol				
t-butanol				