

Report No. (2)

Preparation of A Solid/Liquid Solution+ Serial Dilution

Student Names: Section No:

The objective of the experiment:

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Part one: Dissolving a solid solute in a solvent

Results and calculations:

1. Tabulate your results

m_{NaCl}	V_{solution}

2. Calculate the molarity and the strength of your solution:

Atomic weight : Na = 23 , Cl = 35.5

Part 2: Preparing a solution from another solution by dilution:

Results and calculations:

1. Tabulate your results:

M_{conc}	V_{initial}	M_{dil}	V_{total}

2. Calculate the molarity of your diluted solution

Part 3: Preparing solution by serial dilution:

Results and calculations:

1. Tabulate your results:

Test tube number	M_{conc}	V_{initial}	M_{dil}	V_{total}
1	0.1	1 ml		
2		1 ml		
3		1 ml		

2. Using the dilution equation, calculate the molarity of the three solutions

$M_{\text{dil}(\text{test tube}1)}$:

$M_{\text{dil}(\text{test tube}2)}$:

$M_{\text{dil}(\text{test tube}3)}$: