

# Electrical Transport Properties of [Succinonitrile–Poly(Ethylene Oxide)]–LiX–Co(bpy)<sub>3</sub>(TFSI)<sub>2</sub>–Co(bpy)<sub>3</sub>(TFSI)<sub>3</sub> Solid Redox Mediators, Where X = TFSI or Triflate

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**Abstract.** A fast ion-conducting solid redox mediator (electrolyte) is a basic requirement for all-solid-state dye-sensitized solar cells. In the present paper, we have reported electrical transport properties of [(1-x)SN-xPEO]–LiX–Co(bpy)<sub>3</sub>(TFSI)<sub>2</sub>–Co(bpy)<sub>3</sub>(TFSI)<sub>3</sub> solid electrolytes, where SN stands for a plastic crystal, succinonitrile, PEO for a high-molecular-weight polymeric matrix, poly(ethylene oxide), and X for an anion either TFSI (radius 0.7 nm) or triflate (radius 0.44 nm). The composition,  $x = 0, 0.5$ , and  $1$  in weight fraction yields SN, SN–PEO blend, and PEO-based electrolytes. The plasticizing property of the SN helps to lower the high crystallinity of the PEO. Co(bpy)<sub>3</sub>(TFSI)<sub>2</sub>, and Co(bpy)<sub>3</sub>(TFSI)<sub>3</sub> provide redox species. The results were compared with those of acetonitrile (ACN)-based liquid electrolytes. In comparison to the liquid electrolytes (electrical conductivity,  $\sigma_{25^\circ\text{C}} \sim 1.7 \times 10^{-2} \text{ S cm}^{-1}$ ) the solid ones with TFSI<sup>−</sup> ions exhibited  $\sigma_{25^\circ\text{C}}$  of  $2.1 \times 10^{-3}$  for  $x = 0$ ,  $7.2 \times 10^{-4}$  for  $x = 0.5$ , and  $9.7 \times 10^{-7} \text{ S cm}^{-1}$  for  $x = 1$ . The triflate ions-based solid electrolytes had lower  $\sigma_{25^\circ\text{C}}$ -values, such as  $1.5 \times 10^{-3}$  for  $x = 0$ ,  $3.1 \times 10^{-4}$  for  $x = 0.5$ , and  $6.3 \times 10^{-7} \text{ S cm}^{-1}$  for  $x = 1$ . The electrolytes with  $x = 0$  and  $1$  had Arrhenius-type behavior akin to those of liquid electrolytes; however, the activation energy values were quite high. In contrast, the blend-based electrolytes ( $x = 0.5$ ) had Vogel–Tamman–Fulcher-type behavior and the pseudo-activation energy was less than those of the liquid electrolytes. The differential scanning calorimetry revealed the arrest of the amorphous phase for the blend-based electrolytes because of the SN–PEO–ionic salt interactions. The blend-based electrolytes were thermally stable up to  $125^\circ\text{C}$ , making them superior to the liquid electrolytes for the device application.

**Keywords:** Dye-Sensitized Solar Cells, Redox Mediators, Plastic Crystal, Solid Polymer Electrolyte, Electrical Conductivity.