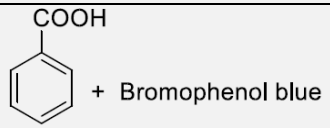
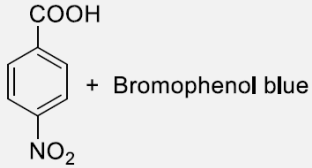
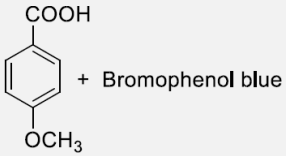
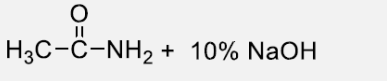
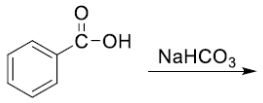
Laboratory Report (109 chem)

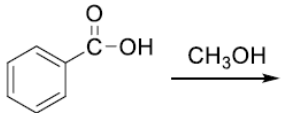
Experiment 8: Carboxylic acid & Their derivatives

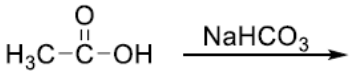
Student Names: …… Section No: ……

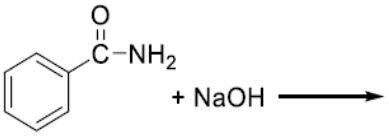
Test	Observation	Result	Chemical equation
Acetic acid + $NaHCO_3$ 1 ml acetic acid +(Δ) Heat for 1 min + 0.5g $NaHCO_3$(solid)			
 + Bromophenol blue	The acid will therefore be stronger if Ar is electron withdrawing than if Ar is electron donating.		——-
 + Bromophenol blue			
 + Bromophenol blue			
Esterification 1- 1ml of acetic acid +1ml Ethanol +2drops Conc.H_2SO_4 + (Δ) Heat for 2 min 2 -Add to test tube contain 10 ml water +Na_2CO_3	Distinctive smell of ester		
 with red litmus paper	red litmus paper changes the colour to blue		

Questions:









Name	class	Functional group	Molecular formula	Structure formula
acetic acid				
Salicylic acid				