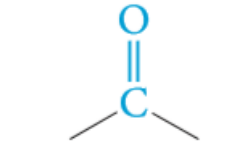
108 Chem

Chapter-6

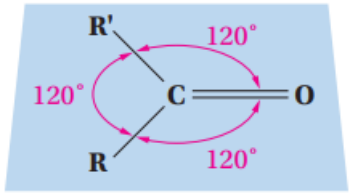
Aldehydes and ketones

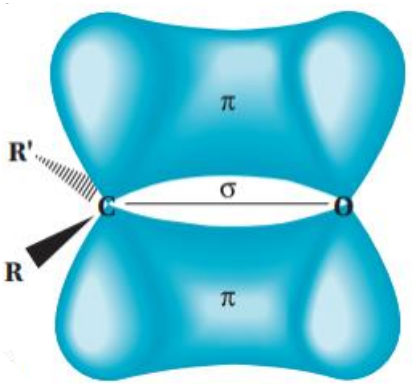
Structural Characteristic of Aldehydes and ketones

- Aldehydes and ketones are characterized by the presence of the carbonyl group.

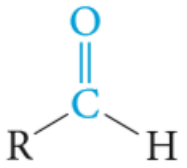


carbonyl group

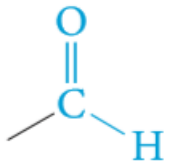




- The functional group of an aldehyde is a carbonyl group bonded least one hydrogen atom attached and the remaining group may be another hydrogen atom or any aliphatic or aromatic organic group.
- The functional group of a ketone is a carbonyl group bonded to two carbon atoms.

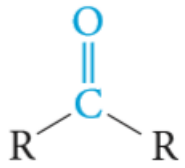


aldehyde



or —CH=O or —CHO

aldehyde group or formyl group



ketone

Nomenclature of aldehydes

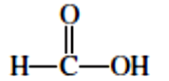
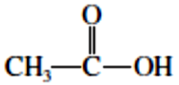
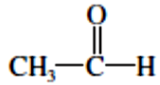
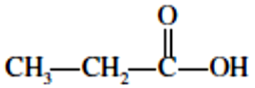
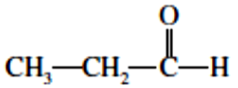
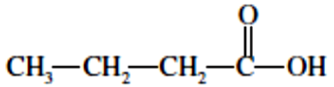
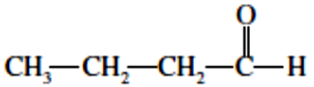
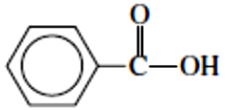
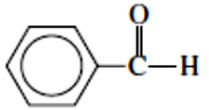
The IUPAC:

- Select the longest chain of carbon atoms that contains the functional group.
- Changing the suffix *-e* of the parent alkane to *-al.*
- Number must start with carbonyl group of an aldehyde as carbon-1.
- For unsaturated aldehydes, the presence of a carbon-carbon double or triple bond is indicated by the infix *-en-* or *-yn-*. As with other molecules with both an infix and a suffix, the location of the group corresponding to the suffix determines the numbering pattern.
- For cyclic molecules in which -CHO is bonded directly to the ring, the molecule is named by adding the suffix *-carbaldehyde* to the name of the ring. The atom of the ring to which the aldehyde group is bonded is numbered 1

Nomenclature of aldehydes

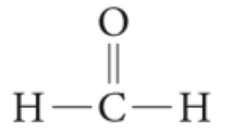
The Common name:

- The common name for an aldehyde is derived from the common name of the corresponding carboxylic acid by dropping the word *acid* and changing the suffix *-ic* or *–oic* to *-aldehyde*.
- In common names carbon atoms near the carbonyl group are often designated using Greek letters (α, β, γ , δ…..) beginning with carbon next the carbonyl group.

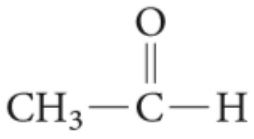
Carboxylic acid	Derivation	Aldehyde
 formic acid (methanoic acid)	*formica*, "ants"	formaldehyde (methanal)
 acetic acid (ethanoic acid)	*acetum*, "sour"	 acetaldehyde (ethanal)
 propionic acid (propanoic acid)	*protos pion*, "first fat"	 propionaldehyde (propanal)
 butyric acid (butanoic acid)	*butyrum*, "butter"	 butyraldehyde (butanal)
 benzoic acid	*gum benzoin*, "blending"	 benzaldehyde

IUPAC name:

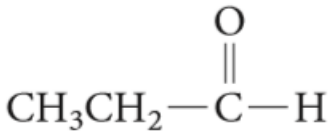
Common name:



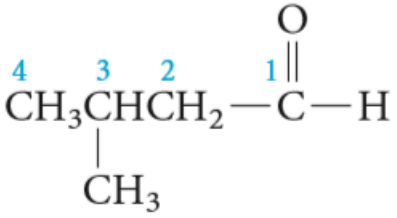
methanal
(formaldehyde)



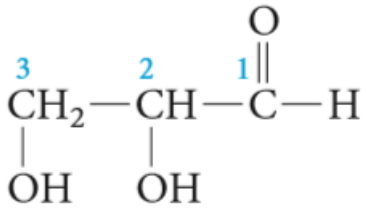
ethanal
(acetaldehyde)



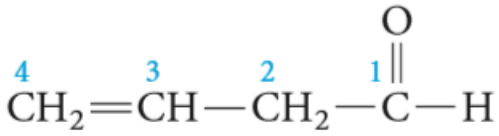
propanal
(propionaldehyde)



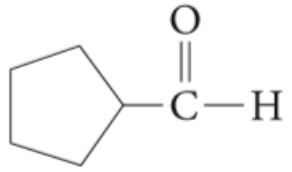
3-methylbutanal



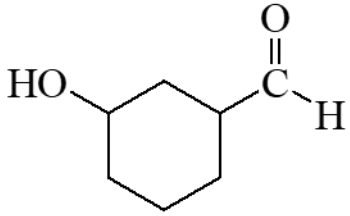
2,3-dihydroxypropanal
(α,β-dihydroxypropionaldehyde)



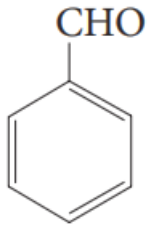
3-butenal



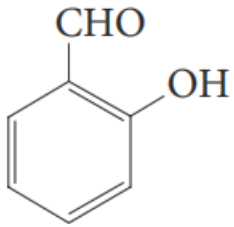
cyclopentanecarbaldehyde



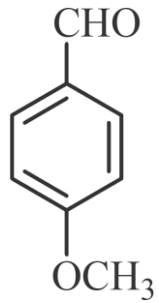
4-Hdroxycyclohexanecarbaldehyde



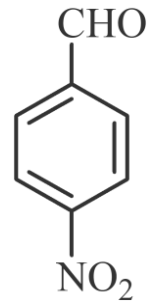
benzaldehyde
(benzenecarbaldehyde)



salicylaldehyde
(2-hydroxybenzenecarbaldehyde)



p-Methoxybenzaldehyde



p-Nitrobenzaldehyde

Nomenclature of Ketones

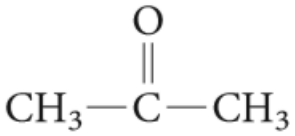
The IUPAC:

- Select the longest chain of carbon atoms that contains the functional group.
- Changing the suffix *-e* of the parent alkane to *-one*
- The parent chain is numbered from the direction that gives the carbonyl carbon the smaller number.
- If a ketone has a second functional group of higher naming priority, the ketone oxygen is indicated by the prefix "*oxo-*."

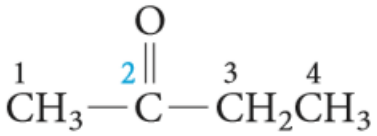
The Common names:

- The common name for ketones are derived by naming the two alkyl or aryl groups bonded to the carbonyl group as separate words followed by the word *ketone*.

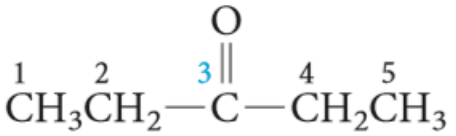
IUPAC name:
Common name:



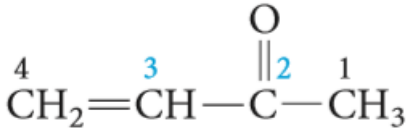
propanone
(acetone)



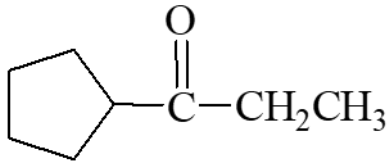
2-butanone
(ethyl methyl ketone)



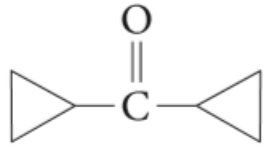
3-pentanone
(diethyl ketone)



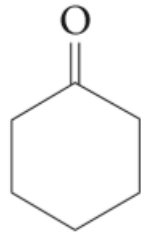
3-buten-2-one
(methyl vinyl ketone)



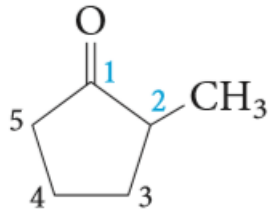
1-cyclopentylpropanone
Ethyl cyclopentyl ketone



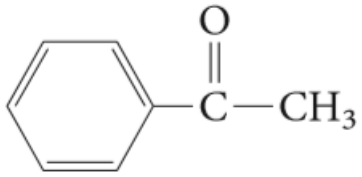
Dicyclopropylmethanone
Dicyclopropyl ketone



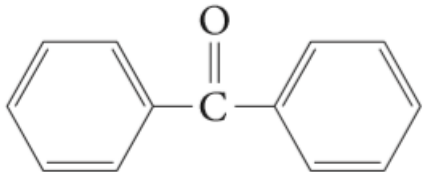
cyclohexanone



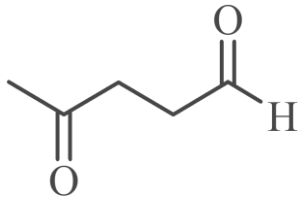
2-methylcyclopentanone



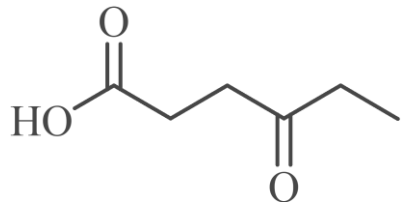
acetophenone
1-phenylethanone
(methyl phenyl ketone)



benzophenone
diphenylmethanone
(diphenyl ketone)



3-oxopentanal



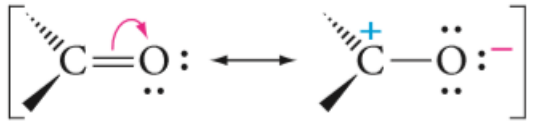
4-Oxohexanoic acid

Priority Order in Nomenclature System

Class	Suffix name	Prefix name
Carboxylic acid	-oic acid	Carboxy
Ester	-oate	Alkoxycarbonyl
Amide	-amide	Amido
Nitrile	-nitrile	Cyano
Aldehyde	-al	Oxo (=O)
Aldehyde	-al	Formyl (CH=O)
Ketone	-one	Oxo (=O)
Alcohol	-ol	Hydroxy
Amine	-amine	Amino
Alkene	-ene	Alkenyl
Alkyne	-yne	Alkynyl
Alkane	-ane	Alkyl
Ether	—	Alkoxy
Alkyl halide	—	Halo

Physical Properties of aldehydes and ketones

Oxygen is much more electronegative than carbon. Therefore, the electrons in the C=O bond are attracted to the oxygen, producing a highly polarized bond.



resonance contributors to the carbonyl group

polarization of the carbonyl group

Boiling points

In general, aldehydes and ketones have **higher boiling points than alkenes** because they are more polar and the dipole–dipole attractive forces between molecules are stronger. But they have **lower boiling points than alcohols** because, unlike alcohols, two carbonyl groups can't form hydrogen bonds to each other.

Solubility:

Aldehydes and ketones can form hydrogen bonds with the protons of OH groups. This makes them more soluble in water than alkenes, but less soluble than alcohols.

	$CH_3CH_2CH{=}CH_2$	$CH_3CH_2CH{=}O$	CH_3CCH_3 (O)	$CH_3CH_2CH_2OH$
	1-Butene	Propanal	Propanone	1-Propanol
bp (1 atm)	−6°C	49°C	56.1°C	97°C
Solubility in water (g/100 mL)	Negligible	20	∞	Miscible in all proportions

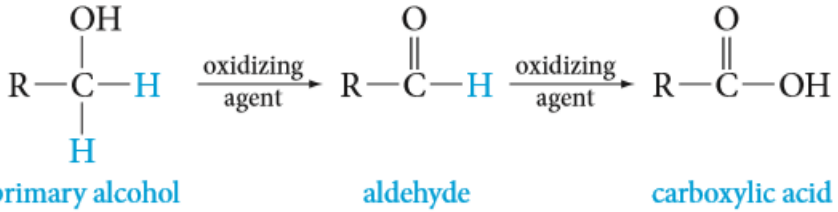
Preparation of Aldehyde and Ketone

1. Oxidation of Primary and Secondary Alcohols

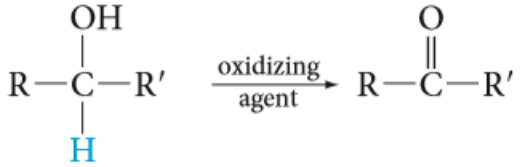
Strong oxidizing agent

Potassium permanganate $KMnO_4$,OH^-/ H_3O^+

Chromic oxide CrO_3 / H_2SO_4 (H_2CrO_4 Jones' reagent)



primary alcohol — aldehyde — carboxylic acid



secondary alcohol — ketone

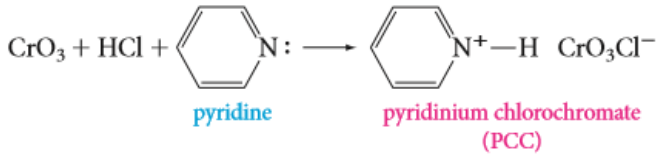
No reaction

Tertiary alcohol

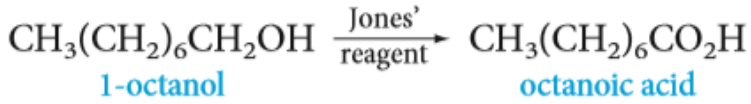
Weak oxidizing agent

Chromic oxide CrO_3 / pyridine

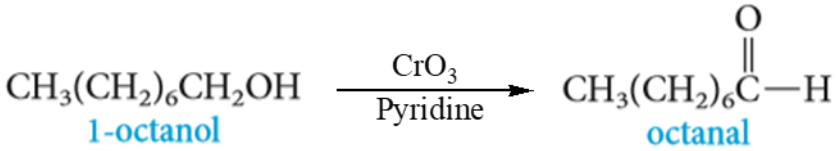
Pyridinium chlorochromate PCC / methylene chloride CH_2Cl_2



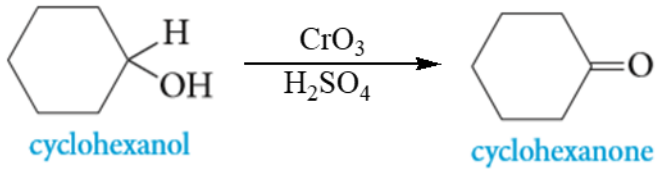
pyridine — pyridinium chlorochromate (PCC)



1-octanol — octanoic acid



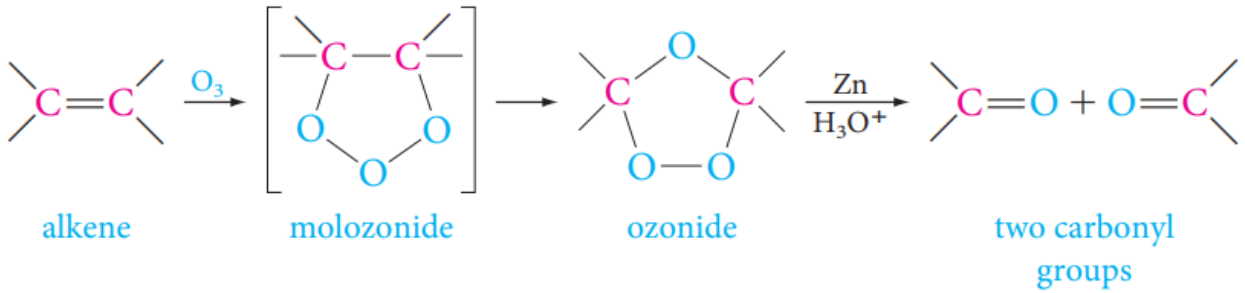
1-octanol — octanal



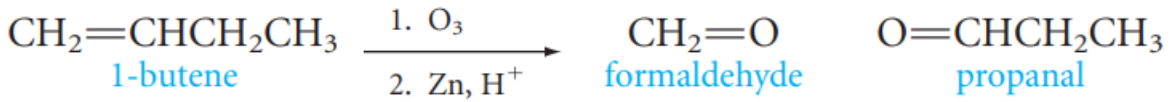
cyclohexanol — cyclohexanone

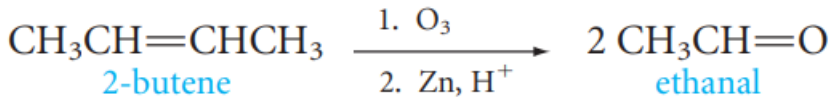
Preparation of Aldehyde and Ketone

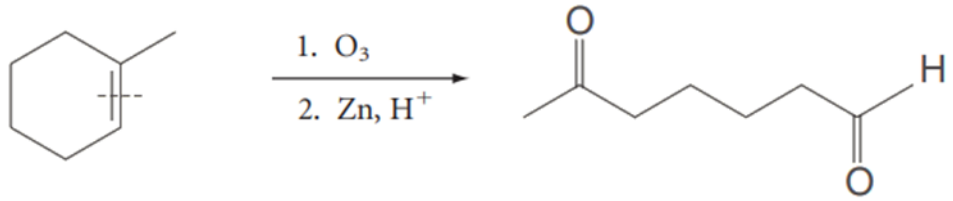
2- Ozonolysis of Alkenes



Examples:



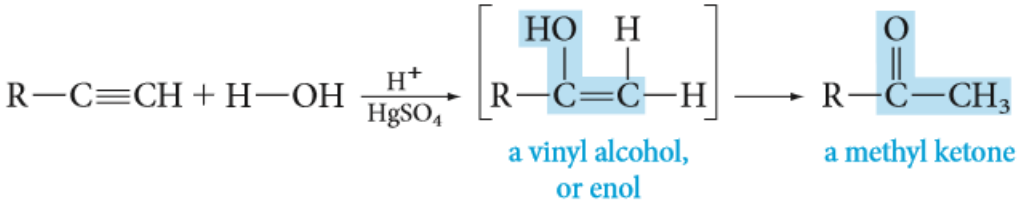


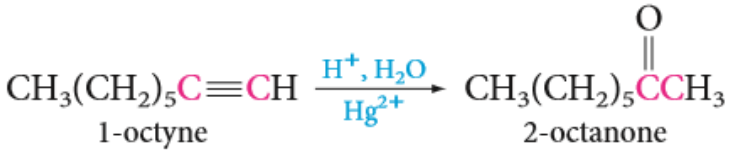


Preparation of Aldehyde and Ketone

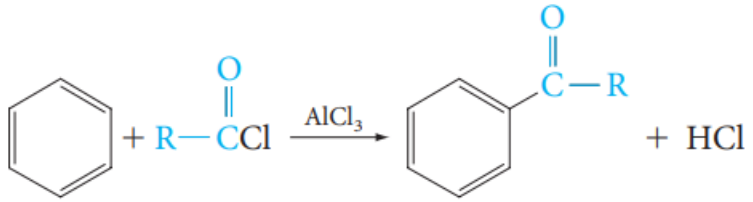
3- Hydration of Alkynes

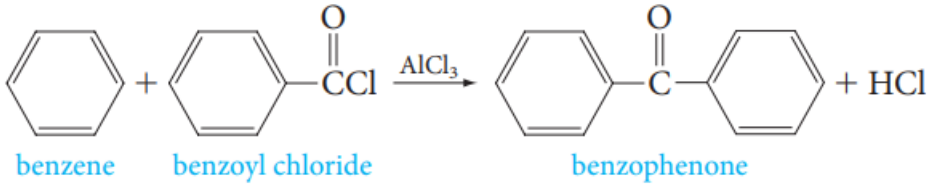
Addition of water to terminal alkynes requires not only an acid catalyst but mercuric ion as well.





4- Friedel–Crafts Acylation: Preparation of Aromatic ketones

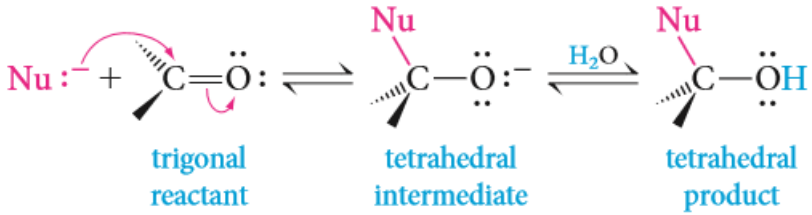




Reaction of Aldehyde and Ketone

1. Nucleophilic addition to carbonyl groups

    - Addition of Grignard Reagents to aldehyde and ketone
    - Reduction of carbonyl group
    - The Addition of Alcohols: Hemiacetals and Acetals
    - Addition of Hydrogen Cyanide: Formation of cynohydrins
    - Addition of Nitrogen Nucleophiles

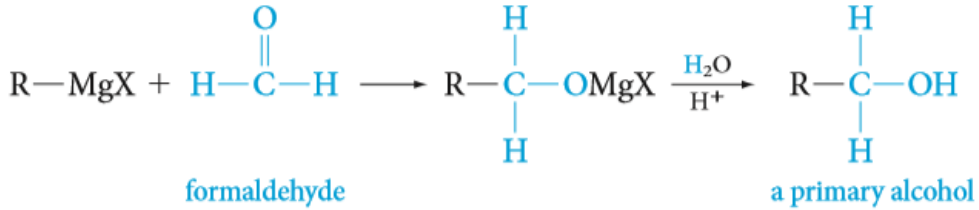


2. Oxidation of Aldehydes

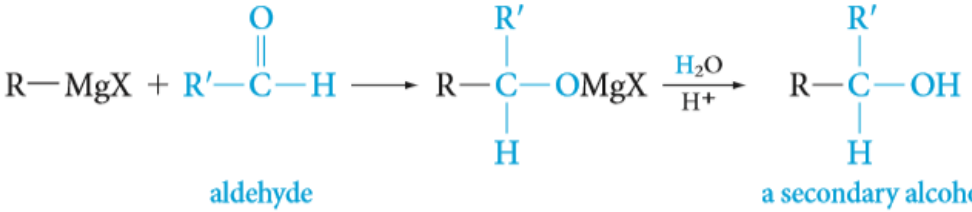
Reaction of Aldehyde and Ketone

1- Nucleophilic addition to carbonyl groups:

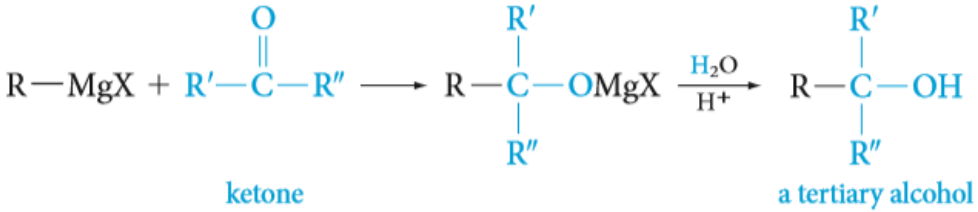
a) Addition of Grignard Reagents to aldehyde and ketone: formation of alcohol

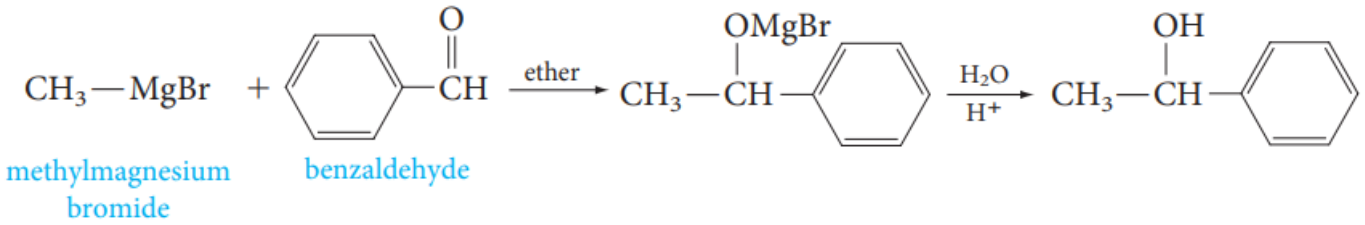


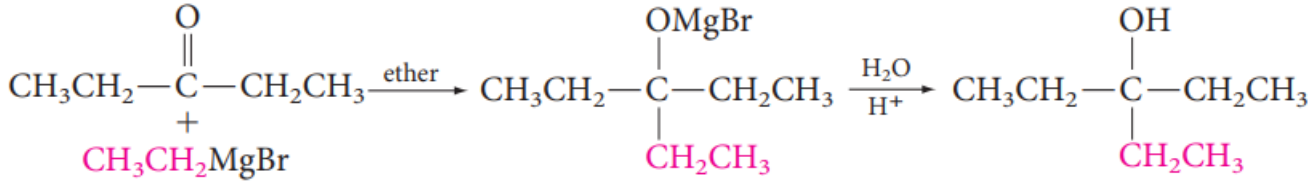
Other aldehydes give secondary alcohols.



Ketones give tertiary alcohols.

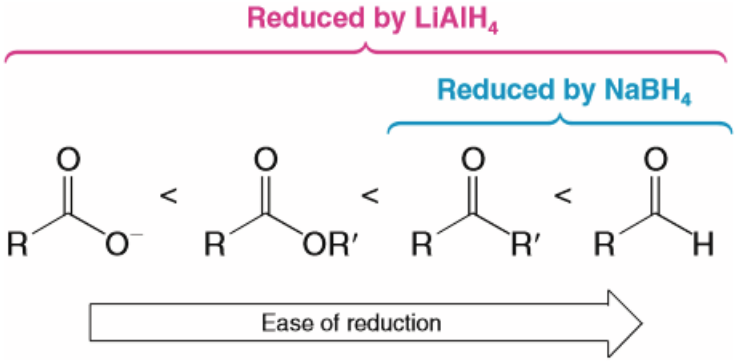




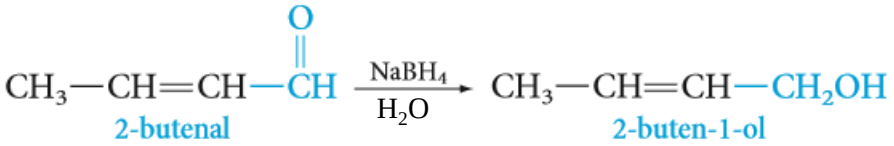


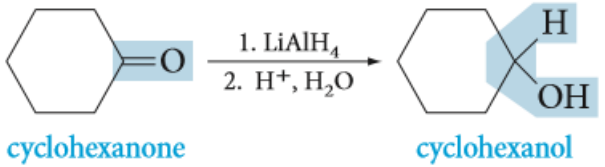
Reaction of Aldehyde and Ketone

b) Reduction of carbonyl group: Addition of metal hydrides (formation of alcohol)



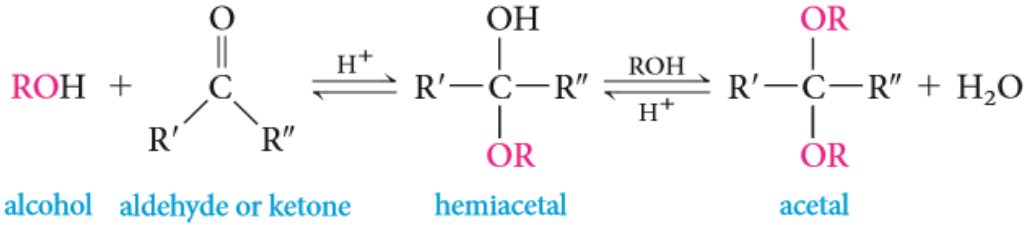
Examples:



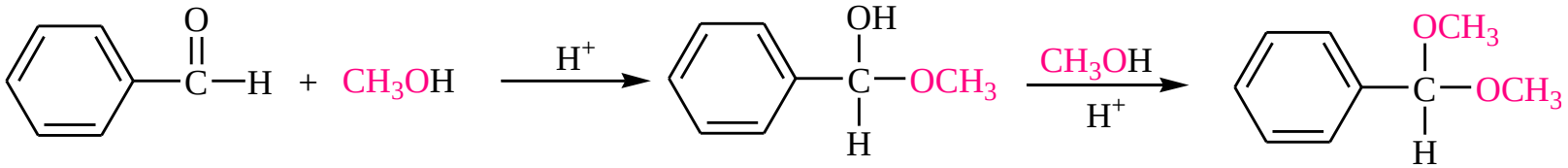


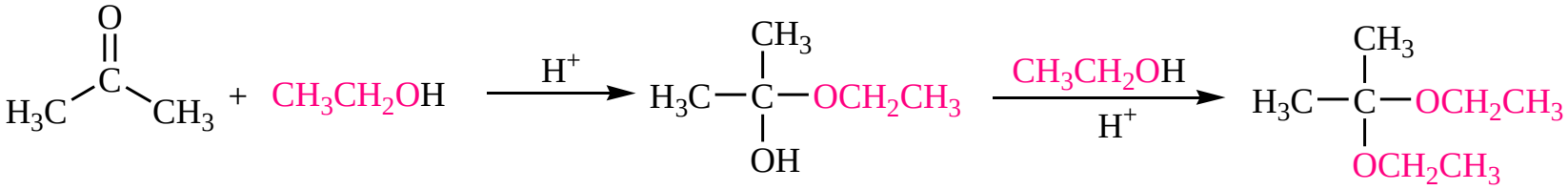
Reaction of Aldehyde and Ketone

c) The Addition of Alcohols: Hemiacetals and Acetals

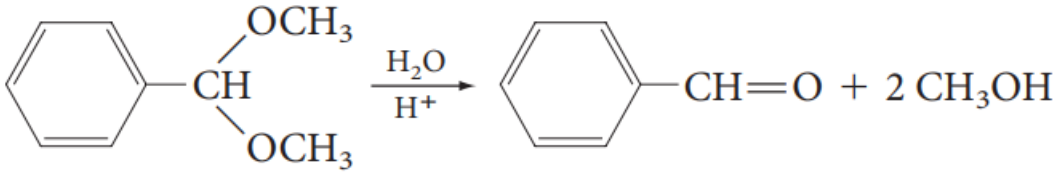


Examples:



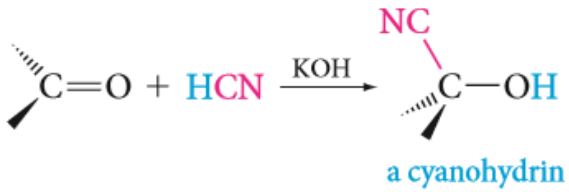


- Acetal can be hydrolyzed to its aldehyde or ketone and alcohol components by treatment with excess water in the presence of an acid catalyst.



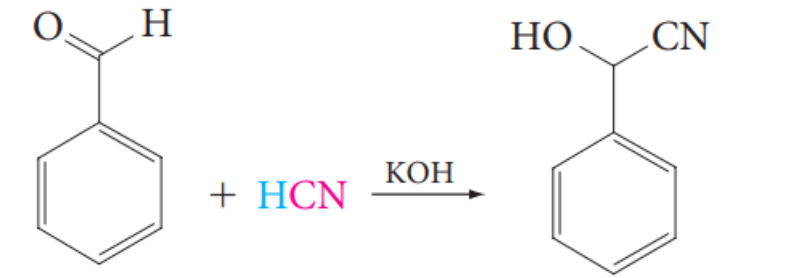
Reaction of Aldehyde and Ketone

d) Addition of Hydrogen Cyanide: Formation of cynohydrins



a cyanohydrin

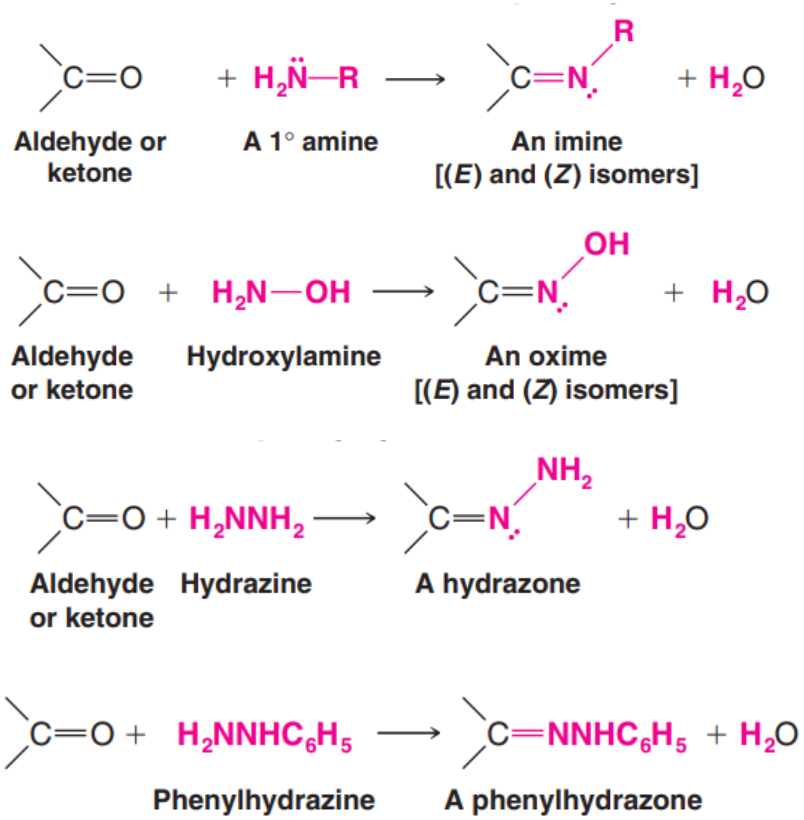
acetone — acetone cyanohydrin



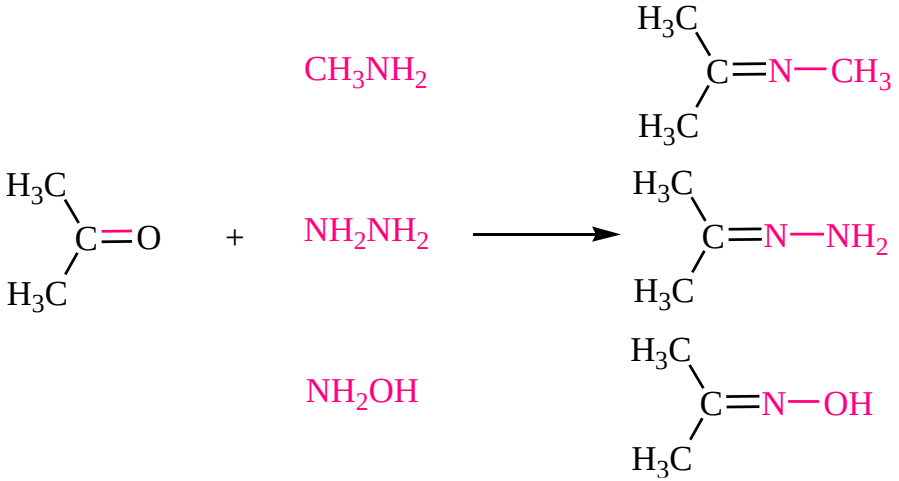
benzaldehyde — benzaldehyde cyanohydrin

Reaction of Aldehyde and Ketone

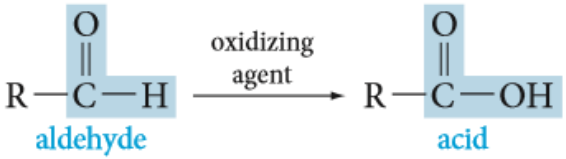
e) Addition of Nitrogen Nucleophiles



Examples:



2- Oxidation of Aldehydes



$$CH_3(CH_2)_5CH{=}O \xrightarrow[\text{(Jones' reagent)}]{CrO_3,\ H^+} CH_3(CH_2)_5CO_2H$$