334 MBIO

Biochemical Instrumentation Techniques

- Lab 2 -

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The pH meter

- pH is the negative logarithm of hydrogen ion activity: pH =
 log [H⁺].
- It is based on the concentration or the activity of hydrogen ions, of either an aqueous solution, or semi-solid substances.





pH Scale and Interpretation

The pH value of a substance is directly related to the ratio of the hydrogen ion and hydroxyl ion concentrations, as follows:

Acidic: pH < 7 (Higher H⁺ concentration)

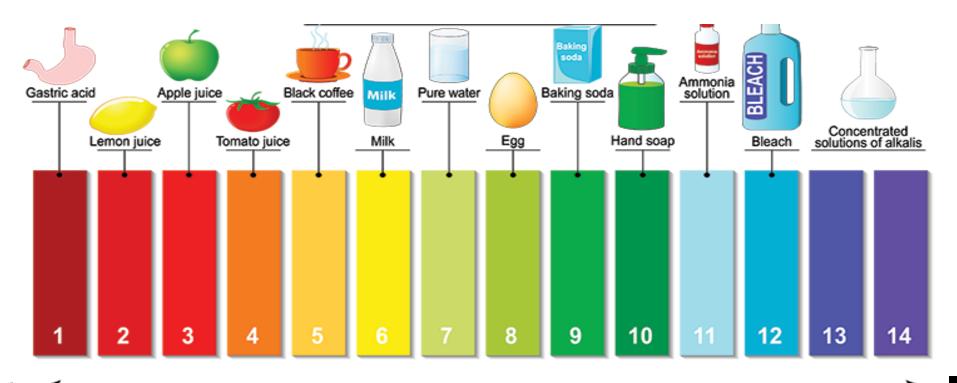
Natural: pH = 7

Alkaline (Base): pH > 7 (Higher OH⁻ concentration)



pH Scale and Interpretation

• It measures the acidity or alkalinity of a solution on a scale of 0 to 14.



pH Scale and Interpretation

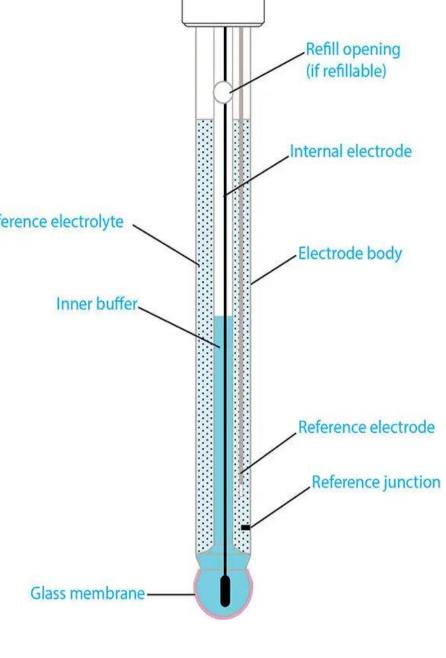
- How to control the pH:
 - Adding acids increases H⁺ ions.
 - Adding bases increases OH⁻ ions.



Components of a pH Meter

• It consists of a pH measuring electrode (pH- Reference electrolyte sensitive electrode), a reference electrode, or liner buffer.

• The pH measuring **electrode** is a hydrogen ion sensitive **glass bulb**.





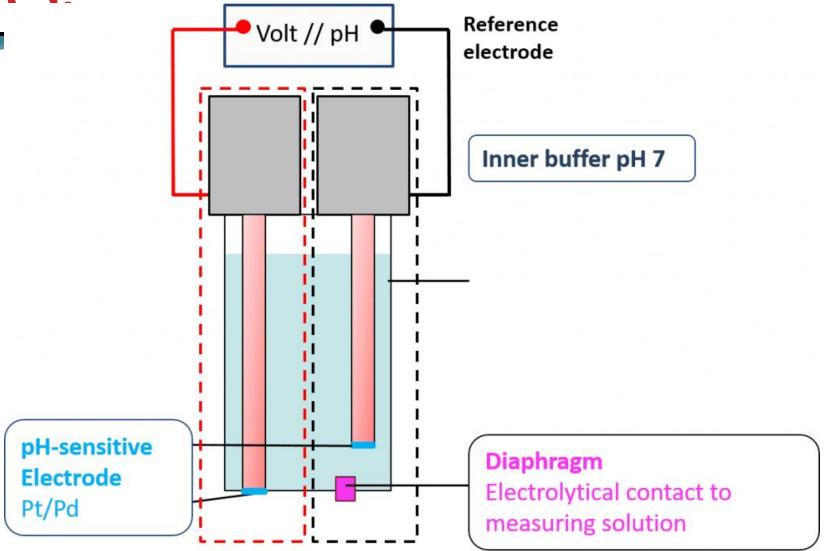
Key Features:

- Electrode Holding Arm: Supports the probe for immersion in the sample.
- **Combination Electrode:** Contains a pH-sensitive glass bulb for detecting H⁺, and reference electrode with stable output.
- Display Panel: Shows the measured pH value.





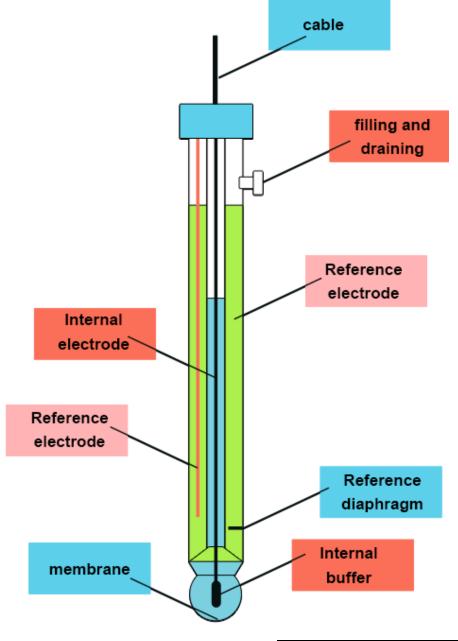
Key Features:





How it works

- The probe contains two electrodes:
 - One in a liquid of fixed pH.
 - One in the sample to detect H⁺ concentration.
- The voltmeter measures the difference in voltage between these electrodes.





How it works

• The meter translates the voltage difference into a pH value, which is displayed on the screen.





Applications of pH Meters

Laboratory Work → Chemical analysis and experiments.

Agriculture → Soil testing for crop optimisation.

Water Quality → Monitoring water supply, swimming pools, and environmental systems

Healthcare → Ensuring that solutions are safe when applied to patients or lethal as sterilant and disinfectants.



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