**312 BCH**

**Exp (2)**

**Title:**

**Objectives:**

**1-**

**2-**

**Introduction:**

**Materials and Method:** As seen in the lab sheet.

**Results and Discussion:**

**Questions:**

**Q1-** A student needed to prepare 1L of a 1M NaCl solution, which of the following methods is more accurate in preparing the solution? Why?

a) Weighing 58.5g of solid NaCl carefully, dissolving it in 300ml of water, and then adding 700ml of water.

b) Weighing 58.5g of solid NaCl carefully, dissolving it in a small volume of water then making the final volume up to 1L by adding water.

**Q2-**List the most important points to be considered when preparing solutions.

**Q3-**A solution was prepared by taking 6ml of a 0.22M solution and then the volume was made up to a final volume of 30ml. What is the concentration of the final solution?

**Q4-**How would you prepare 80ml of a 1:25 dilution of a 2.1M KCl solution?