

Laboratory Report

Experiment 7: Aldehydes and Ketones

Name:.....

Tube no.	Test	Observation	Conclusion
1	Oxidation of Aldehyde (Formaldehyde or Acetaldehyde) by Potassium permanganate. 1 mL of Aldehyde + 1 drop of KMnO₄		
2	Oxidation of Ketone (Acetone) by Potassium permanganate. 1 mL of Ketone + 1 drop of KMnO₄		
3	Distinguishing Test Oxidation of Aldehyde by weak oxidizing agent (Tollen's reagent) 0.5 mL of AgNO₃ + 2-3 drops of 5% NaOH + 4-5 drops of NH₄OH, + 0.5 mL of Aldehyde + heating for a few seconds		

Questions:

- 1) How can we identify aldehydes and ketones?
- 2) Write the equation represents the reduction of benzaldehyde.
- 3) Write the equation represents the oxidation of p-methoxy benzaldehyde.
- 4) Write the equation represents the oxidation of Acetaldehyde by Tollens reagent.