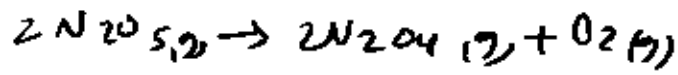


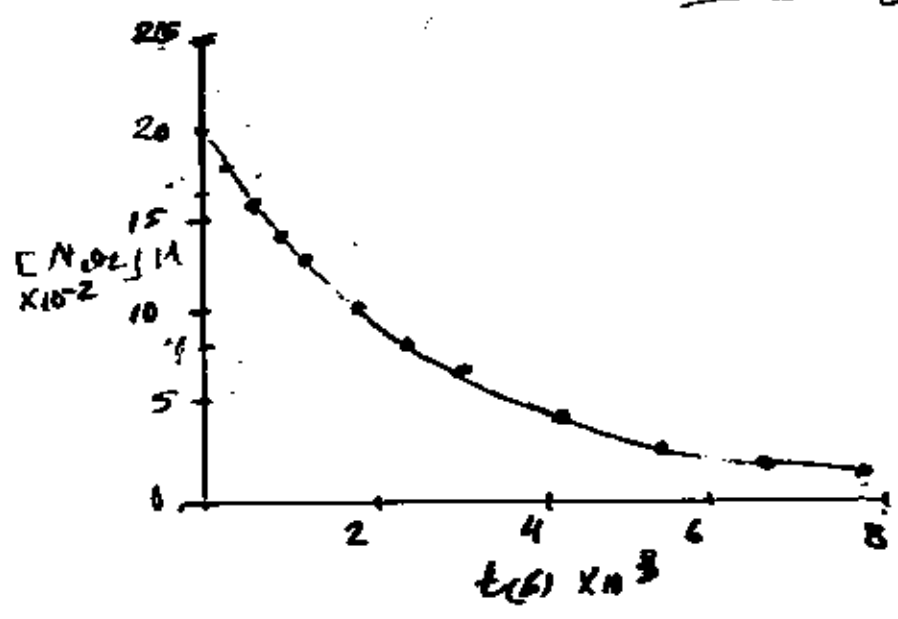
(10)



تفاعل من الدرجة 1

$t(s) \times 10^2$	0	3	6	9	12	18	24	30	42	54
$[N_2O_5] M \times 10^{-2}$	20	18	16.1	14.4	13	10.4	8.4	6.8	4.4	2.8

Rate $\frac{d[N_2O_5]}{dt}$ في وقت معين



$t(s) \times 10^2$	Rate $M \times s^{-1}$
0	7.29
300	6.46
6	5.80
9	5.21
12	4.69
18	3.79
24	2.04
30	2.44
42	1.59
54	1.03

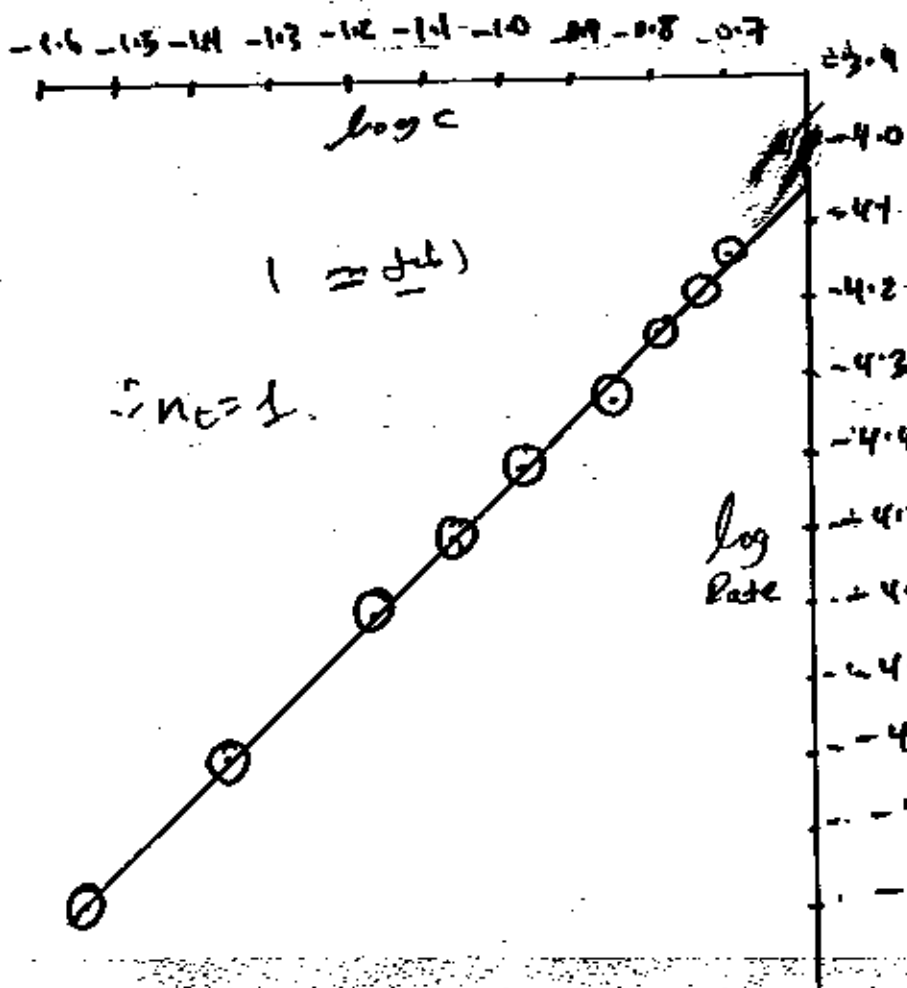
$Rate = k [N_2O_5]^1$

$s^{-1} = \text{Rate}, 3.65 \times 10^{-4} = \frac{Rate}{[N_2O_5]}$

$R \propto [A]$

أو k ثابت

الخط المستقيم



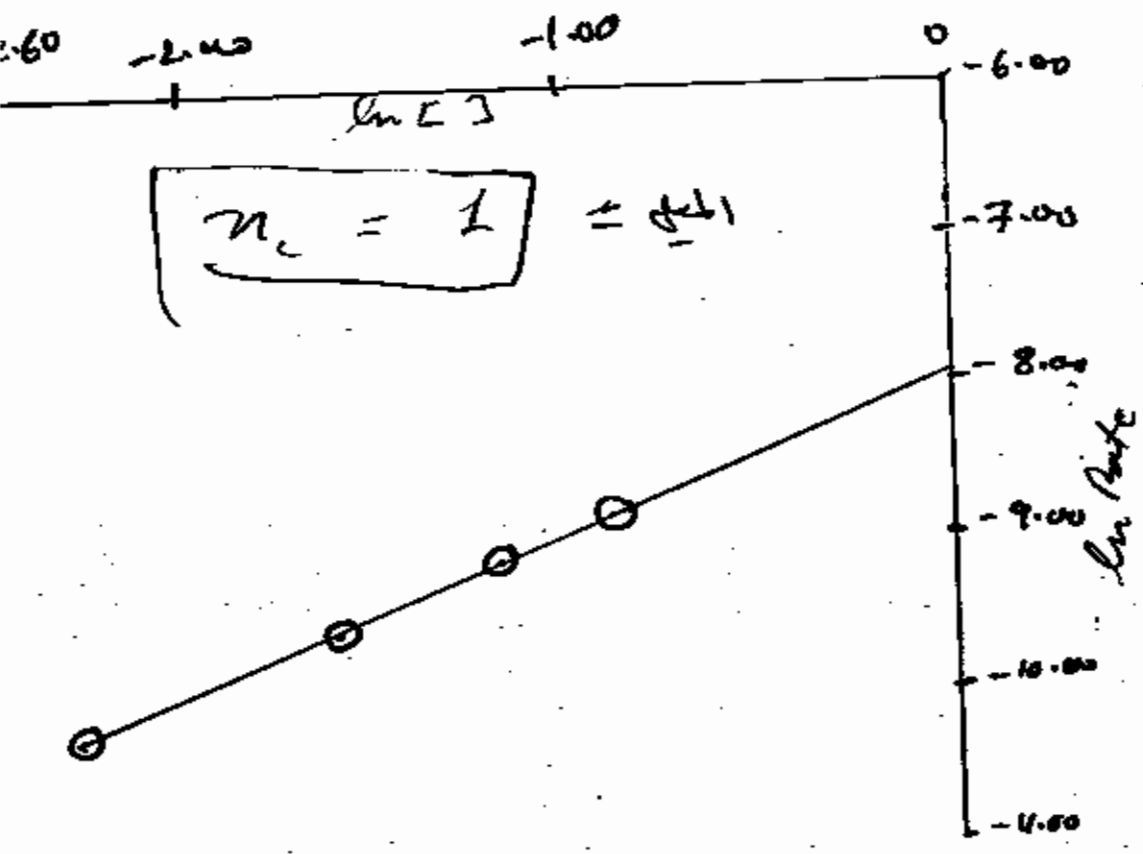
$(1 \approx \text{slope})$
 $\therefore n = 1$

$\log C$	$\log Rate$
-1.70	-4.14
-1.74	-4.19
-1.79	-4.24
-1.85	-4.28
-1.89	-4.33
-1.98	-4.42
-1.06	-4.52
-1.17	-4.61
-1.36	-4.80
-1.55	-4.99

$\log k = \frac{\log Rate}{[C]}$
 $k = 3.11 \times 10^{-5}$

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العمق (م)	ΣN2027 M	Rate (م/س) $(M s^{-1})$	ln L	ln Rate
1	0.100	3.62×10^{-5}	-2.30	-10.23
2	0.200	7.29×10^{-5}	-1.61	-9.53
3	0.300	10.94×10^{-5}	-1.20	-9.12
4	0.400	14.53×10^{-5}	-0.92	-8.87



$ln k = -7.90$
 $= -7.90$
 $k = 3.71 \times 10^{-4} \text{ M s}^{-1}$