

() : 1427/11/30 :

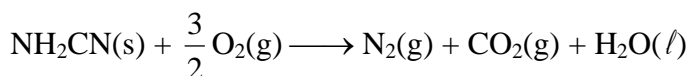


H = 1, C = 12, N = 14, O = 16, F = 19, Ne = 20.2, Na = 23, Cl = 35.5, K = 39.1, :
S = 32, Cr = 52, Zn = 65.4, Br = 80, I = 127
 $N_A = 6.02 \times 10^{23}$, $R = 0.0821 \text{ atm L mol}^{-1} \text{ K}^{-1} = 8.314 \text{ J mol}^{-1} \text{ K}^{-1}$:

(2500 J) (1500 J) (1

: (J) .
+1000 (-4000 (-1000 (+400 (

: (2

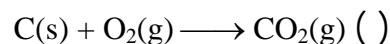
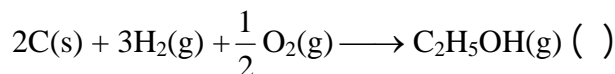
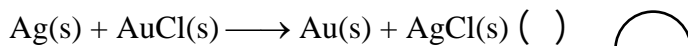
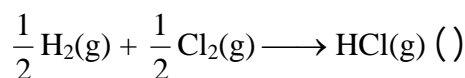


ΔH .(25°C) (-742.7 kJ)

:(kJ)

(-741.5 (+742.7 (-743.9 (

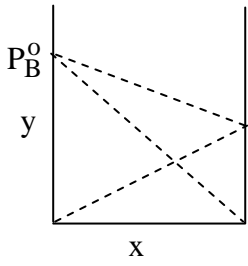
: (3



(0.02 g) (4

: (350 torr) .(700 torr)

0.04 g (2.4×10^{-3} g (0.01 g (1.3×10^{-5} g (



(5)

(y) (x) ()
 (y) (x) ()
 (y) (x) ()
 (y) (x) ()



(25°C)

(950 ml)

(0.2 g)

(6)

(g/mol)

(21.7 torr)

74 (120 (180 (60 (



: _____

(7)

()
 ()
 ()
 ()



: (8)

()
 ()



(50 mmHg)

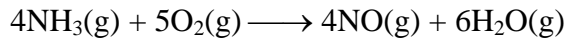
(9)

: (0.1)

5 (55 (32 (45 (



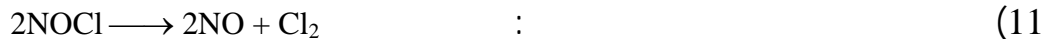
(10)



0.24 M s⁻¹ (NH₃)

0.36 (0.30 (0.24 (0.16 (





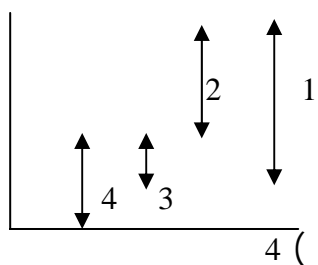
[NOCl] M	(M s^{-1})
0.30	3.60×10^{-9}
0.60	1.44×10^{-8}
0.90	3.24×10^{-8}

() () () ()



(M) (0.052 M) (4.75 s) N_2O_5 $t_{1/2}$
 (10 s)

0.003 () 0.006 () 0.025 () 0.012 ()



(ΔH)

(13)

($5.2 \times 10^{-4} \text{ s}^{-1}$) (300 K) ($2.6 \times 10^{-8} \text{ s}^{-1}$) (14)

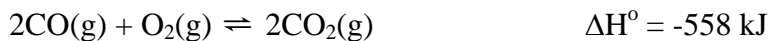
(400 K)

.200 kJ

$t_{1/2}$

(15)

(16)



.T

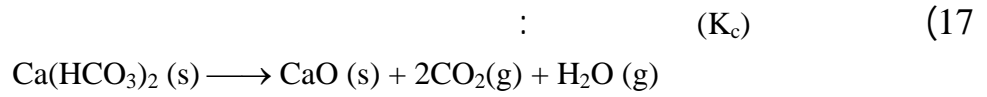
CO

O₂

CO₂

.CO₂

CO



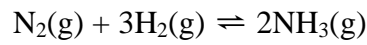
$$K_c = \frac{[\text{CaO}][\text{CO}_2]^2[\text{H}_2\text{O}]}{[\text{Ca(HCO}_3)_2]} \quad ($$

$$K_c = [\text{CO}_2]^2 [\text{H}_2\text{O}] ($$

$$K_c = 2[\text{CO}_2][\text{H}_2\text{O}] ($$

$$K_c = \frac{[\text{Ca(HCO}_3)_2]}{[\text{CaO}][\text{CO}_2]^2[\text{H}_2\text{O}]} ($$

$$:(472^\circ\text{C}) \quad (K_c = 0.105) \quad (18)$$



$$1.55 \times 10^2 ($$

$$3.73 \times 10^2 ($$

$$2.81 \times 10^{-5} ($$

$$5.25 \times 10^{-4} ($$

$$: \quad K_p \quad ($$

$$: \quad (19)$$

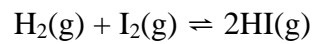


$$($$

$$($$

$$\text{H}_2 \quad ($$

$$($$



$$($$

$$($$

$$($$

$$($$